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of a compound is the percent by mass
of each of the elements in a

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compound. 2. Divide the mass of each element in the sample by the mass of the sample. Then multiply each quotient by 100. 3.c.

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$R = 0.0821 \text{ liter}\cdot\text{atm}/\text{mol}\cdot\text{K}$ ($P = \text{atm}$, V

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= L and T = K) $R = 8.3145 \text{ J/mol}\cdot\text{K}$
(Pressure x Volume is energy, T = K)
 $R = 8.2057 \text{ m}^3 \cdot \text{atm/mol}\cdot\text{K}$ (P = atm, V
= cubic meters and T = K) $R =$
 $62.3637 \text{ L}\cdot\text{Torr/mol}\cdot\text{K}$ or
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